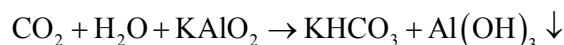
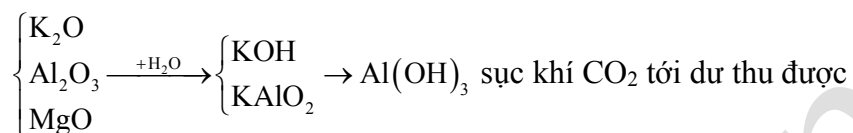


Đáp án

1-C	2-D	3-C	4-A	5-A	6-B	7-C	8-D	9-D	10-D
11-A	12-B	13-C	14-A	15-B	16-D	17-A	18-C	19-D	20-B
21-A	22-A	23-A	24-C	25-A	26-B	27-C	28-C	29-B	30-D
31-A	32-C	33-D	34-C	35-D	36-B	37-D	38-A	39-A	40-C

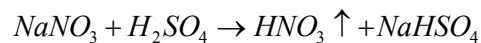
LỜI GIẢI CHI TIẾT

Câu 20: Đáp án B

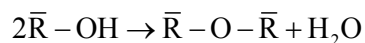


Câu 21: Đáp án A

Câu 22: Đáp án A



Câu 23: Đáp án A



$$30g \qquad 25,5g$$

$$\Rightarrow m_{H_2O} = 4,5g \Rightarrow n_{H_2O} = 0,25mol$$

$$\Rightarrow \bar{R} = 43(C_3H_7)$$

$$\Rightarrow \bar{n} = 3 \Rightarrow A$$

Câu 24: Đáp án C

$$n_e = \frac{I.t}{F} = 0,013$$

$$n_{AgNO_3} = 0,004mol$$

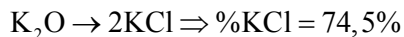
$$\Rightarrow m = 0,432g$$

Câu 25: Đáp án A

Câu 26: Đáp án B

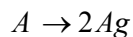
$$m_{K_2O} = 47g$$

$$\Rightarrow n_{K_2O} = 0,5mol \Rightarrow n_{KCl} = 0,5.2 = 1mol$$



Câu 27: Đáp án C

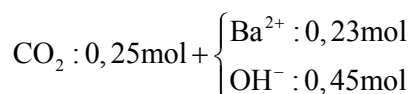
Giả sử ra hai bạc



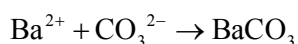
$$0,2 \leftarrow 0,4mol$$

$$M_A = 29(CHO) \Rightarrow C$$

Câu 28: Đáp án C

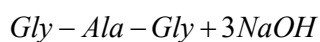


$$* = \frac{n_{OH^-}}{n_{CO_2}} = 1,8 \Rightarrow n_{CO_3^{2-}} = n_{OH^-} - n_{CO_2} = 0,2mol$$



$$\Rightarrow m = 39,40g$$

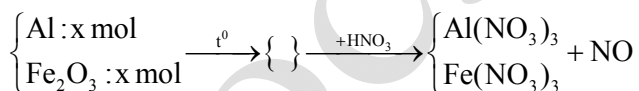
Câu 29: Đáp án B



$$0,15 \qquad \qquad 0,45mol$$

$$\Rightarrow m = 18g$$

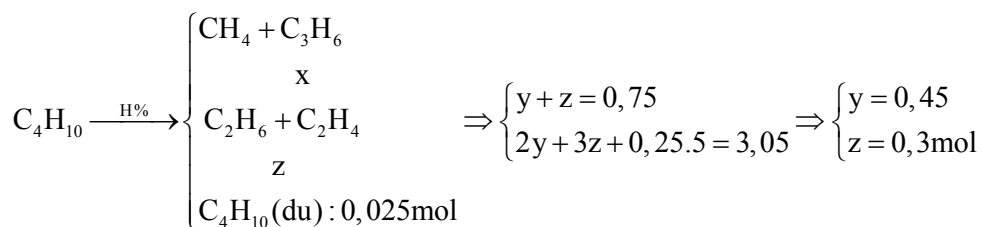
Câu 30: Đáp án D



$$\Rightarrow m = 3,74$$

Câu 31: Đáp án A

$$x = 1mol$$



$$\%CH_4 = 23,45\%$$

Câu 32: Đáp án C

$$A(\dots; 0,8) \Rightarrow n_{Al^{3+}} = 0,8 \text{ mol}$$

$$B(x; 0,2)$$

$$n \downarrow = 4n_{Al^{3+}} - n_{OH^-} \Leftrightarrow 0,2 = 4 \cdot 0,8 - x \Rightarrow x = 3,0$$

Câu 33: Đáp án D

$$n_{este} = 0,6 \text{ mol} \Rightarrow R-COONa: 0,6 \text{ mol}$$

$$n_{Na_2CO_3} = 0,5 \text{ mol}; \Rightarrow \text{Este của phenol: } 0,04 \text{ mol}$$

$$n_{CO_2} = 2,5 \text{ mol}$$

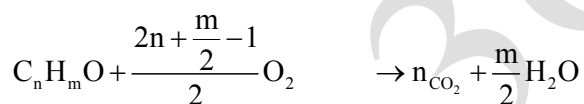
$$\text{BTNT C: } 0,6(n+1) + 0,4(m+6) = 3$$

$$\Leftrightarrow 0,6n + 0,4m = 0$$

$$\Rightarrow \begin{cases} n = 0 \\ m = 0 \end{cases} \Rightarrow \begin{cases} \text{HCOONa: } 0,6 \\ \text{C}_6\text{H}_5\text{ONa: } 0,4 \end{cases} \text{ mol}$$

$$\Rightarrow m_{H_2O} = 23,4 \text{ g}$$

Câu 34: Đáp án C



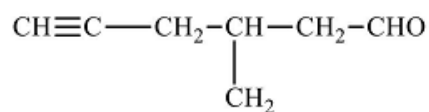
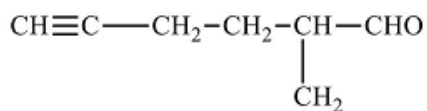
$$0,04 \rightarrow 0,02 \left(2n + \frac{m}{2} - 1 \right) \quad 0,04n \quad 0,02m$$

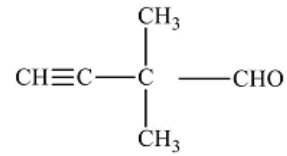
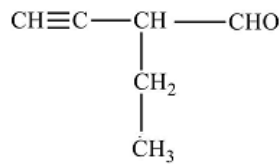
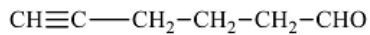
$$\Rightarrow 0,44 = 0,04n + 0,02m + 0,34 - 0,02 \left(2n + \frac{m}{2} - 1 \right)$$

$$\Leftrightarrow 0,1 = 0,01m + 0,02 \Rightarrow m = 8 (C_nH_8O)$$

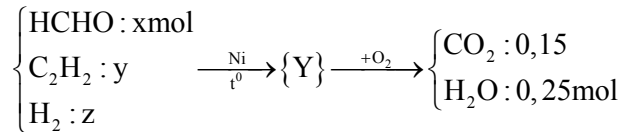
X \rightarrow 2Ag $\downarrow \Rightarrow \downarrow$ của an-1-in

$$\frac{2n + 2 - 8}{2} = 3 \Rightarrow n = 6 (C_6H_8O)$$





Câu 35: Đáp án D



*BTNT "O"

$$n_{\text{HCHO}} + 2n_{\text{O}_2} = 2n_{\text{CO}_2} + n_{\text{H}_2\text{O}} \Rightarrow x = 0,05 \text{ mol}$$

*BTNT : C

$$n_{\text{HCHO}} + 2n_{\text{C}_2\text{H}_2} = n_{\text{CO}_2}$$

$$\Rightarrow n_{\text{C}_2\text{H}_2} = 0,05 \text{ mol} = y$$

*BTNT : H

$$2n_{\text{HCHO}} + 2n_{\text{C}_2\text{H}_2} + 2n_{\text{H}_2} = 2n_{\text{H}_2\text{O}}$$

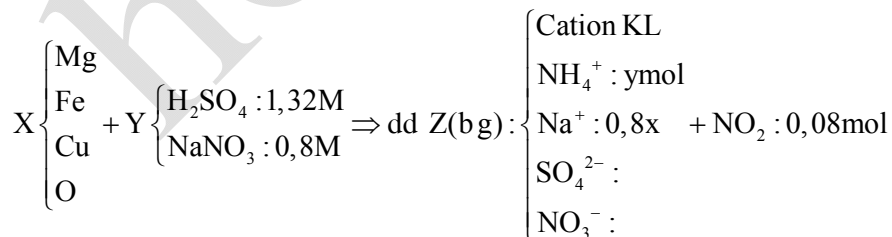
$$\Rightarrow z = 0,15 \text{ mol}$$

$$\% \text{C}_2\text{H}_2 = \frac{m_{\text{C}_2\text{H}_2} \cdot 100\%}{\sum m} = 41,94\%$$

Câu 36: Đáp án B

Câu 37: Đáp án D

$$a(\text{g}) \begin{cases} m_{\text{O}} = 0,2a(\text{g}) \\ m_{\text{KL}} = 0,8a(\text{g}) \end{cases} \mid b = 3,66a$$



$$Z + \text{KOH}(1,22 \text{ mol})$$

$$\text{Đặt } \begin{cases} V_{\text{ddY}} = x \text{ mol} \\ n_{\text{NH}_4^+} = y \text{ mol} \end{cases}$$

$$n_{H^+} = 4n_{NO} + 10n_{NH_4^+} = 0,32 + 10y$$

$$BT"H": 1,32x \cdot 2 = 0,32 + 10y + \frac{0,2a}{16} \cdot 2$$

$$\Leftrightarrow 2,64x = 0,32 + 10y + 0,025a(1)$$

$$BT"N": 0,8x = y + n_{NO_3^-} + 0,08$$

$$\Rightarrow n_{NO_3^-} = 0,8x - y - 0,08(\text{mol})$$

$$*m_Z = b(\text{g})$$

$$3,66a = 0,8a + 18y + 23 \cdot 0,8x + 96 \cdot 1,32x + 62(0,8x - y - 0,08)$$

$$\Leftrightarrow -2,86a - 44y + 194,72x = 4,96(2)$$

$$dd Z + KOH \rightarrow \begin{cases} K^+ : 1,22\text{mol} \\ Na^+ : 0,8x(\text{mol}) \\ SO_4^{2-} : 1,32x(\text{mol}) \\ NO_3^- : 0,8x - y - 0,08 \end{cases}$$

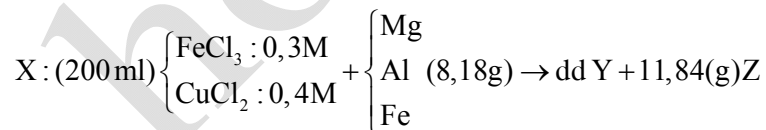
$$BTDT : 1,22 + 0,8x = 2,64x + 0,8x - y - 0,08(3)$$

$$\text{Từ (1), (2), (3) ta có: } \begin{cases} x = 0,51 \\ y = 0,02\text{mol} \\ a = 32\text{g} \end{cases}$$

$$\Rightarrow b = 3,66a = 117,12(\text{g}) \Rightarrow D$$

Câu 38: Đáp án A

Câu 39: Đáp án A



$$dd Y \begin{cases} Fe^{2+} : 0,04(\text{mol}) \\ Mg^{2+} : x(\text{mol}) \\ Al^{3+} : y(\text{mol}) \\ Cl^- : 0,34(\text{mol}) \end{cases} \Rightarrow BTDT : 2x + 3y = 0,26$$

$$BTKL : m_{KL} + m_{m'} = m_Y + m_Z \Rightarrow m_Y = 16,89$$

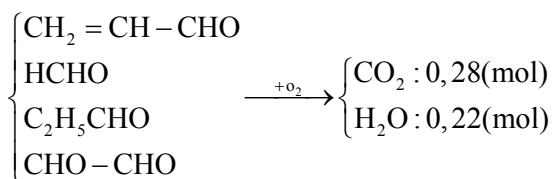
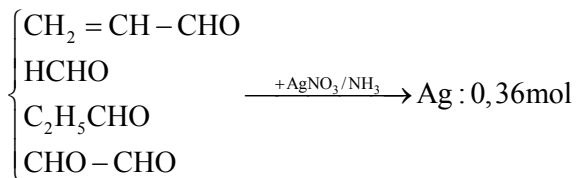
$$\Rightarrow 24x + 27y = 2,58$$

$$\Rightarrow \begin{cases} x = 0,04\text{mol} \\ y = 0,06\text{mol} \end{cases}$$

$$n_{\text{NaOH}} = 2n_{\text{Fe}^{2+}} + 2n_{\text{Mg}^{2+}} + 4n_{\text{Al}^{3+}} = 0,4(\text{mol})$$

$$\Rightarrow m_{\text{NaOH}} = 16(\text{g})$$

Câu 40: Đáp án C



$$* - \text{CHO}(0,18) \rightarrow 2\text{Ag}(0,36) \Rightarrow n_{\text{O}} < 0,18\text{mol} \Rightarrow m_{\text{O}} < 2,88$$

$$m = m_{\text{C}} + m_{\text{H}} + m_{\text{O}} = 3,8 + m_{\text{O}} < 3,8 + 2,88 = 6,686(\text{g})$$

$$\Rightarrow \text{C} : 6,36(\text{g})$$