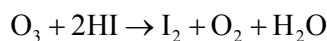
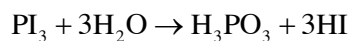
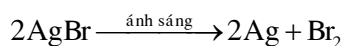
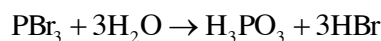
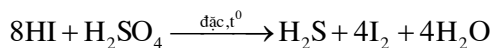
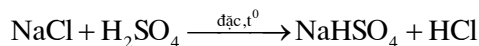
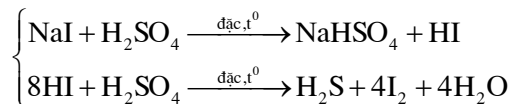
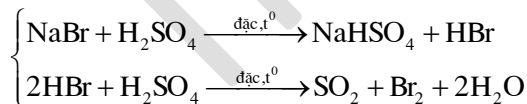
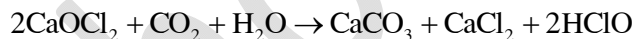
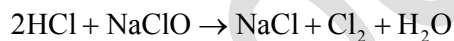
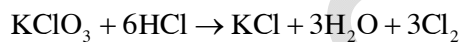
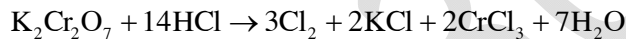
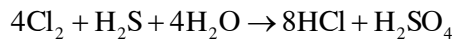
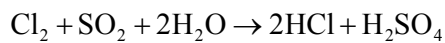
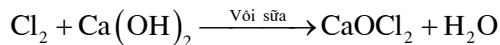
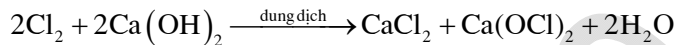
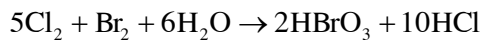
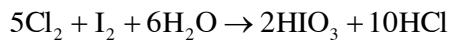
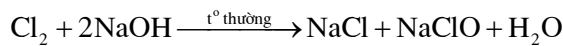
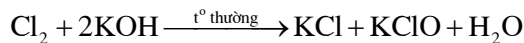
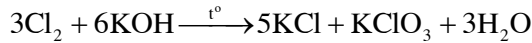
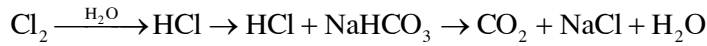
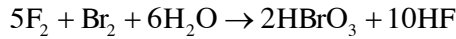
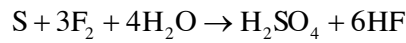
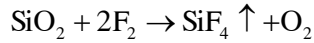
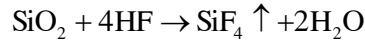
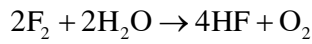
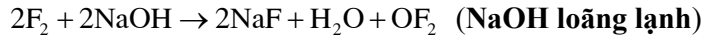
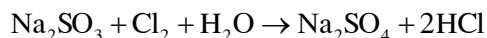
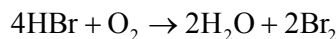
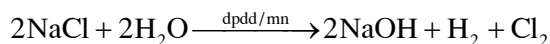
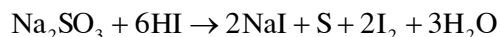
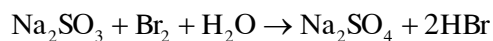
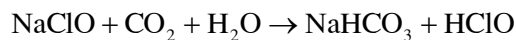


CHƯƠNG NHỮNG VẤN ĐỀ LÝ THUYẾT HÓA HỌC THPT TỔNG HỢP

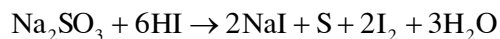
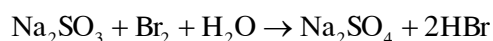
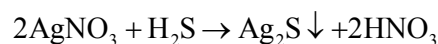
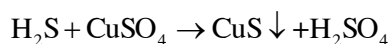
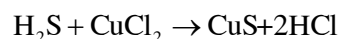
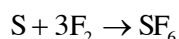
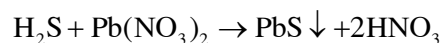
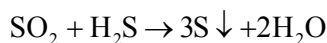
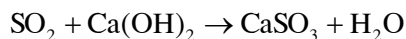
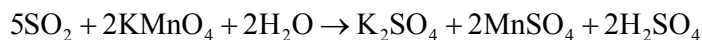
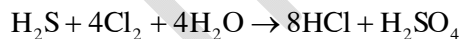
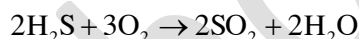
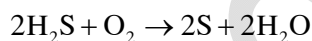
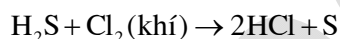
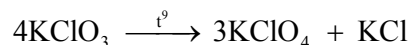
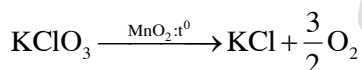
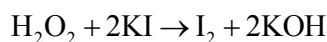
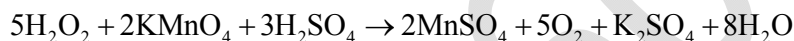
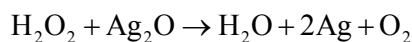
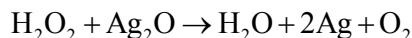
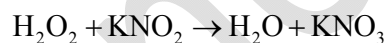
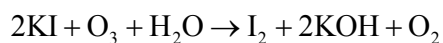
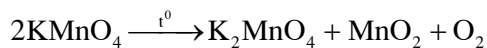
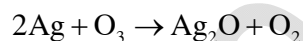
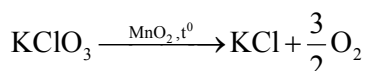
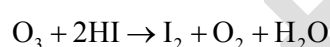
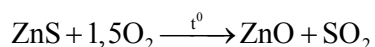
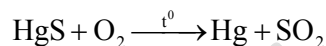
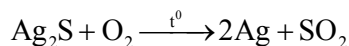
1.1 Những phản ứng trọng tâm cần nhớ

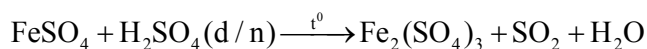
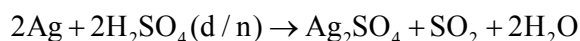
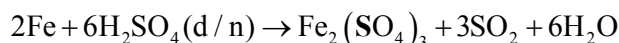
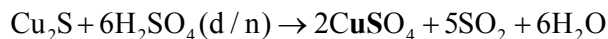
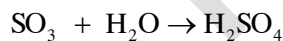
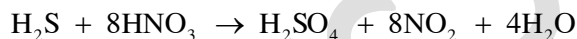
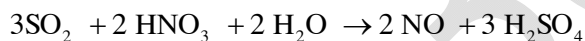
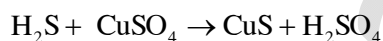
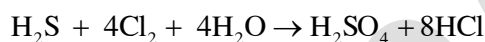
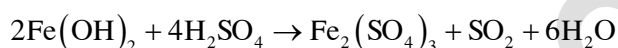
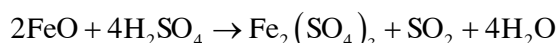
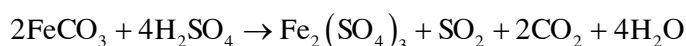
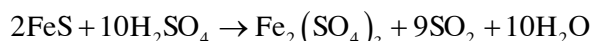
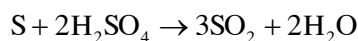
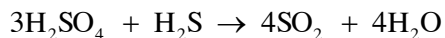
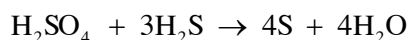
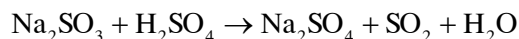
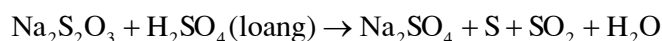
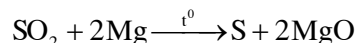
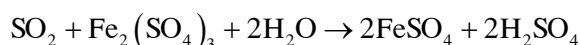
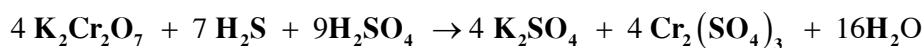
CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI HALOGEN



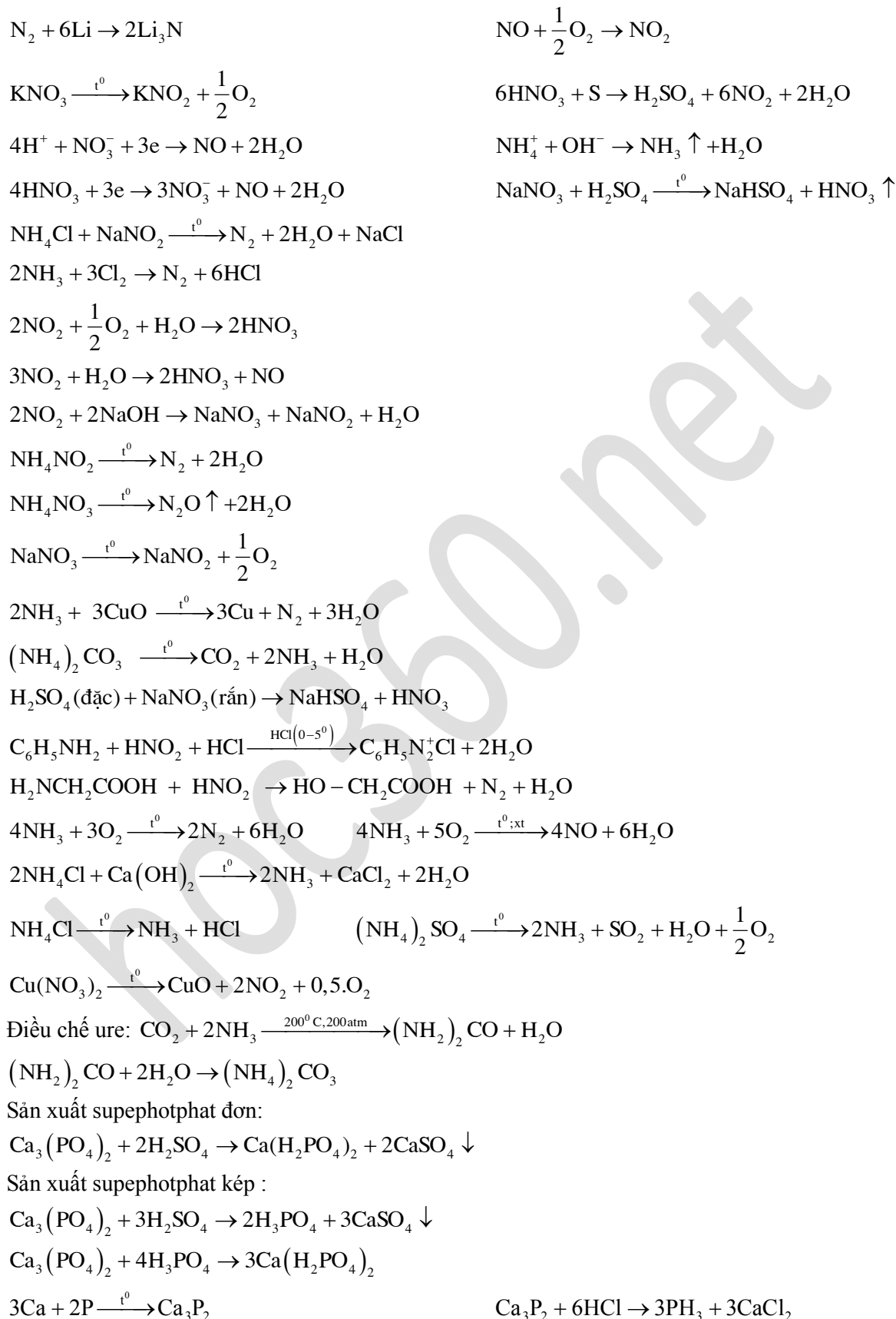


CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI OXI – LƯU HUỖNH





CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI NITƠ – PHOTPHO



Điều chế P trong công nghiệp : $\text{Ca}_3(\text{PO}_4)_2 + 3\text{SiO}_2 + 5\text{C} \xrightarrow{t^0} 3\text{CaSiO}_3 + 2\text{P} + 5\text{CO}$

$2\text{P} + 5\text{H}_2\text{SO}_4(\text{d/n}) \xrightarrow{t^0} 2\text{H}_3\text{PO}_4 + 5\text{SO}_2 + 2\text{H}_2\text{O}$

Phân amophot là hỗn hợp : $\text{NH}_4\text{H}_2\text{PO}_4$ và $(\text{NH}_4)_2\text{HPO}_4$

Phân nitrophotka là hỗn hợp KNO_3 và $(\text{NH}_4)_2\text{HPO}_4$

CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI CACBON – SILIC

$\text{C} + \text{H}_2\text{O} \rightarrow \text{CO} + \text{H}_2$ $\text{C} + 2\text{H}_2\text{O} \rightarrow \text{CO}_2 + 2\text{H}_2$

$\text{CO}_2 + \text{Na}_2\text{SiO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SiO}_3 \downarrow + \text{Na}_2\text{CO}_3$

$\text{HCOOH} \xrightarrow{\text{H}_2\text{SO}_4/\text{đac}} \text{CO} + \text{H}_2\text{O}$

$2\text{Mg} + \text{CO}_2 \rightarrow 2\text{MgO} + \text{C}$

$2\text{Mg} + \text{SO}_2 \rightarrow 2\text{MgO} + \text{S}$

$2\text{H}^+ + \text{CO}_3^{2-} \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

$\text{H}^+ + \text{HCO}_3^- \rightarrow \text{CO}_2 \uparrow + \text{H}_2\text{O}$

$\text{OH}^- + \text{HCO}_3^- \rightarrow \text{CO}_3^{2-} + \text{H}_2\text{O}$

$\text{CO}_2 + \text{Na}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow 2\text{NaHCO}_3$

$\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}$

$\text{C} + 2\text{CuO} \rightarrow \text{CO}_2 + 2\text{Cu}$

$\text{C} + 4\text{HNO}_3 \xrightarrow{t^0} \text{CO}_2 + 4\text{NO}_2 + 2\text{H}_2\text{O}$

$\text{C} + 2\text{H}_2\text{SO}_4 \xrightarrow{t^0} \text{CO}_2 + 2\text{SO}_2 + 2\text{H}_2\text{O}$

$3\text{C} + 2\text{KClO}_3 \xrightarrow{t^0} 2\text{KCl} + 3\text{CO}_2$

$\text{C} + \text{CO}_2 \xrightarrow{t^0} 2\text{CO}$

$\text{Mg} + \text{Si} \xrightarrow{t^0} \text{Mg}_2\text{Si}$

$\text{SiO}_2 + 2\text{NaOH}(\text{nóng chảy}) \xrightarrow{t^0} \text{Na}_2\text{SiO}_3 + \text{H}_2\text{O}$

$\text{SiO}_2 + \text{Na}_2\text{CO}_3(\text{nóng chảy}) \xrightarrow{t^0} \text{Na}_2\text{SiO}_3 + \text{CO}_2$

$\text{SiO}_2 + 2\text{C} \rightarrow \text{Si} + 2\text{CO}$

$\text{SiO}_2 + 2\text{Mg} \xrightarrow{t^0} \text{Si} + 2\text{MgO}$

$\text{Si} + 2\text{NaOH} + \text{H}_2\text{O} \rightarrow \text{Na}_2\text{SiO}_3 + 2\text{H}_2 \uparrow$

$\text{Na}_2\text{SiO}_3 + 2\text{HCl} \rightarrow \text{H}_2\text{SiO}_3 \downarrow + 2\text{NaCl}$

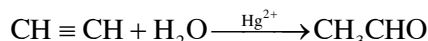
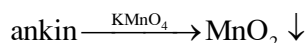
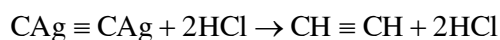
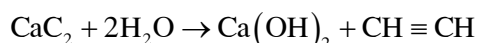
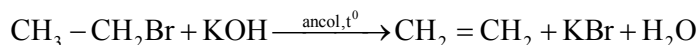
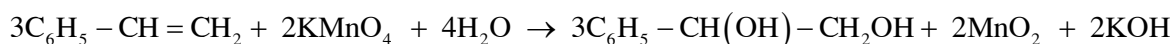
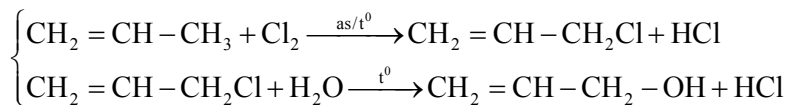
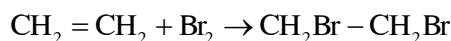
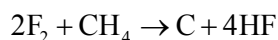
CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI HIDROCACBON

$2\text{CH}_4 \xrightarrow{1500^\circ\text{C}, \text{lnn}} \text{CH} \equiv \text{CH} + 3\text{H}_2$

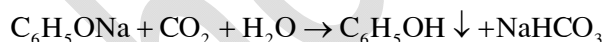
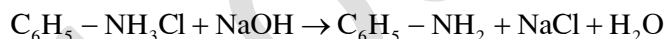
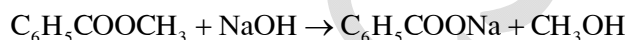
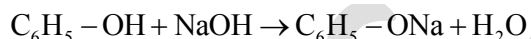
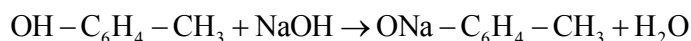
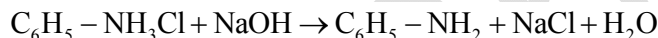
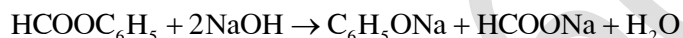
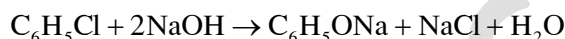
$\text{C}_4\text{H}_{10} \xrightarrow{\text{cracking}} \text{CH}_4 + \text{C}_3\text{H}_6$

$\text{Al}_4\text{C}_3 + 12\text{H}_2\text{O} \rightarrow 4\text{Al}(\text{OH})_3 \downarrow + 3\text{CH}_4$

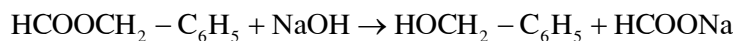
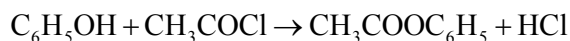
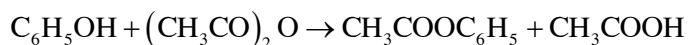
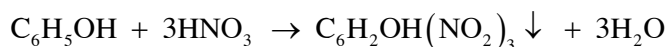
$\text{CH}_3\text{COONa} + \text{NaOH} \xrightarrow{\text{CaO}, t^0} \text{CH}_4 \uparrow + \text{Na}_2\text{CO}_3$

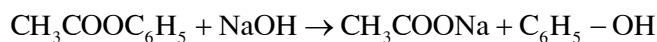


CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI CHẤT CÓ VÒNG BENZEN

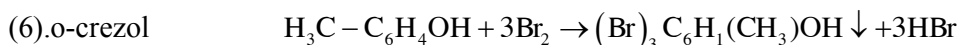
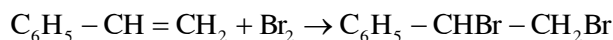
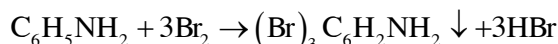
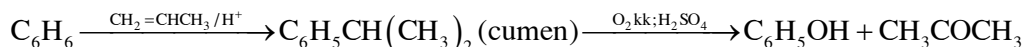


(Trắng)

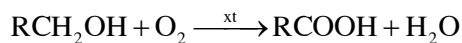
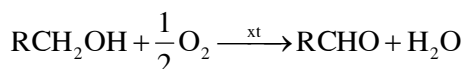
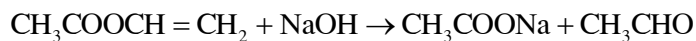
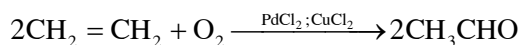
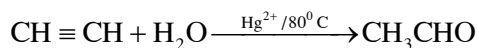
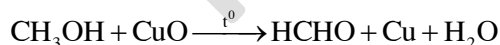
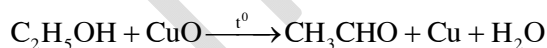
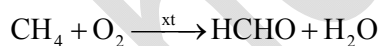
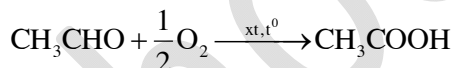
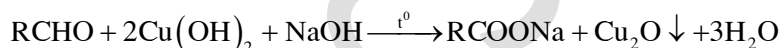
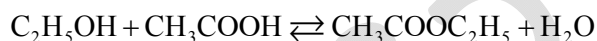
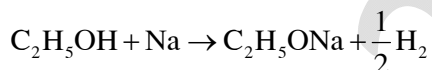
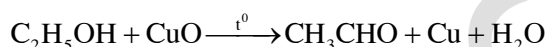
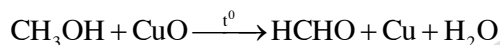
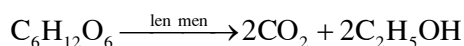
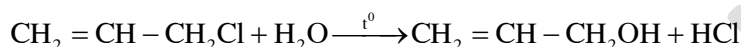
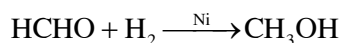
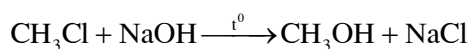
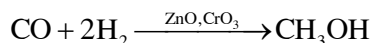


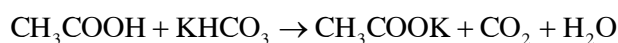
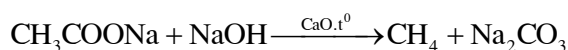
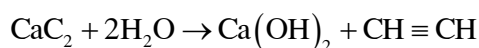
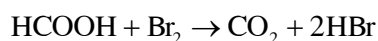
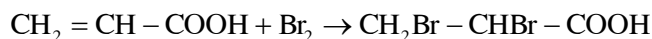
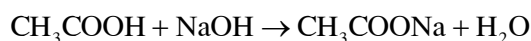
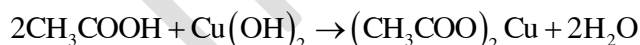
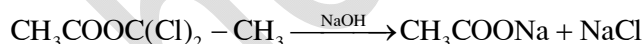
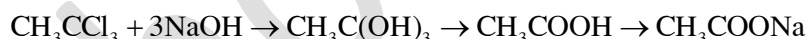
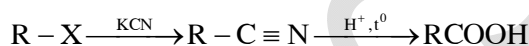
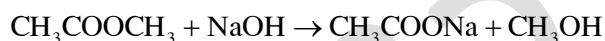
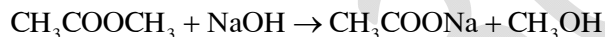
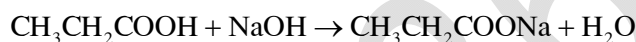
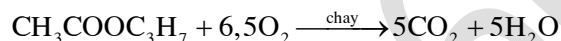
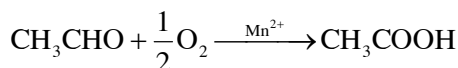
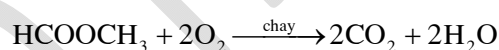
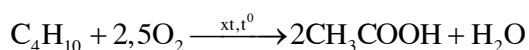
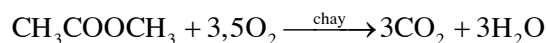
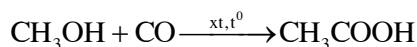
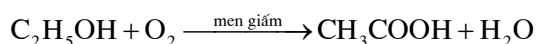
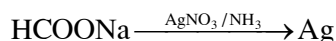
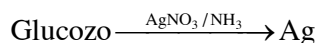
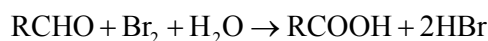
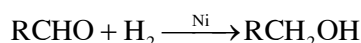
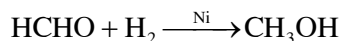
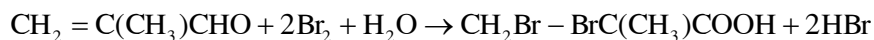
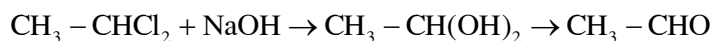


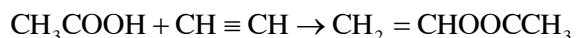
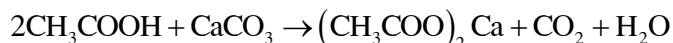
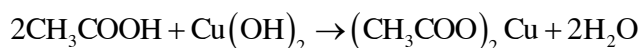
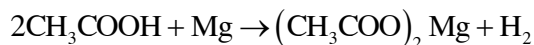
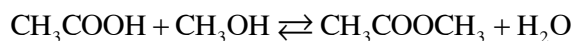
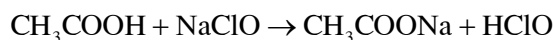
Điều chế phenol và axeton



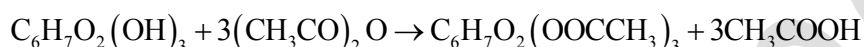
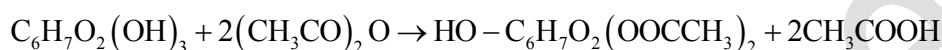
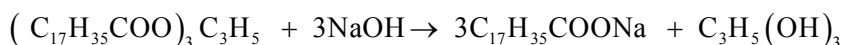
CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI ANCOL - ANDEHIT - AXIT - ESTE



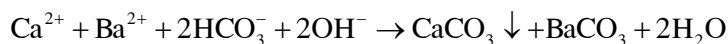
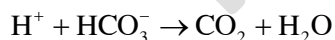
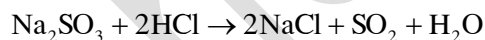
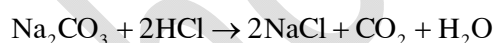
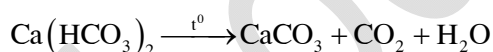
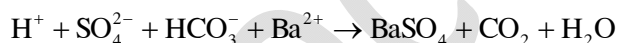
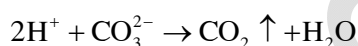
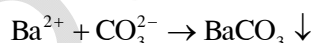
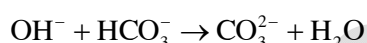
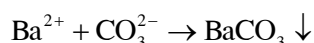
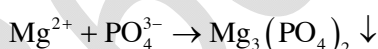
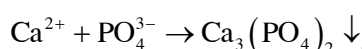
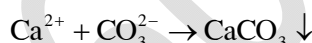
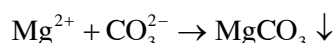
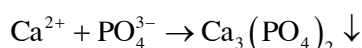
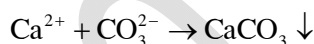
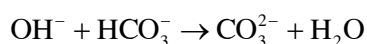
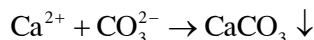




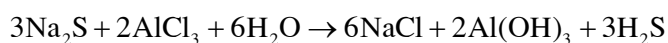
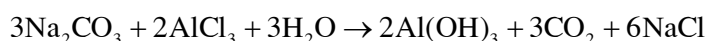
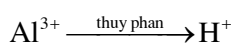
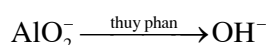
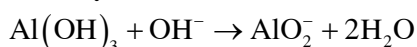
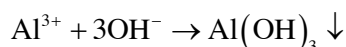
Nhớ 4 loại axit béo quan trọng sau :

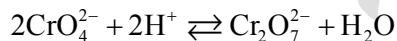
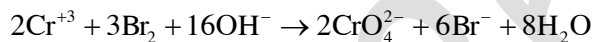
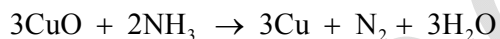
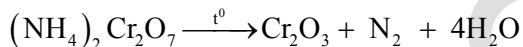
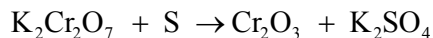
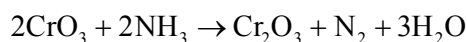
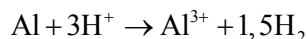
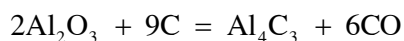
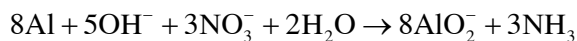
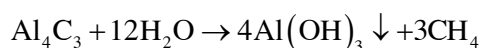
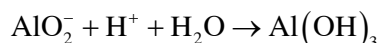
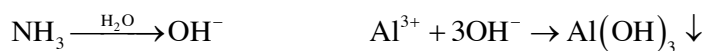
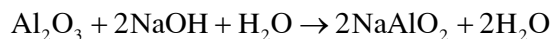
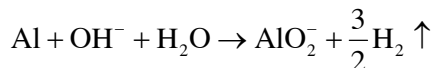
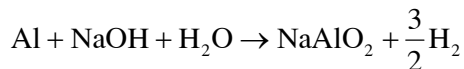
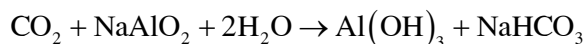


CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI KIM LOẠI KIỀM THỔ



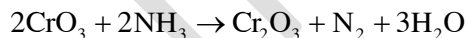
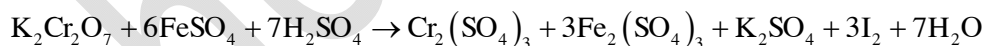
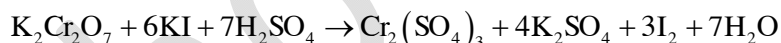
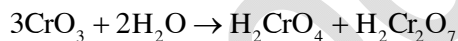
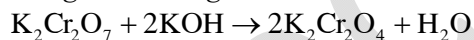
CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI NHÔM – CROM





(màu vàng) (màu da cam)

Trong môi trường axit Zn dễ khử muối Cr^{+3} về Cr^{+2} . $\text{Zn} + 2\text{Cr}^{+3} \rightarrow 2\text{Cr}^{+2} + \text{Zn}^{+2}$



CÁC PHẢN ỨNG QUAN TRỌNG LIÊN QUAN TỚI SẮT

